of Houston and Hsieh. The value of *a* determined by this method using the estimate of 3 percent for the displacements to be accounted for, is about 5×10^{-12} cm. The hypothesis of a positively infinite potential decreases *a* to about 4×10^{-12} cm.

The values of *a* are surprisingly large but could not be reduced to the classical electron radius of 2×10^{-13} cm by

any modification of the experimental results which did not affect the order of magnitude of the recently discovered discrepancies.

> E. C. Kemble R. D. Present

Harvard University, December 2, 1933.

The Infrared Spectrum of Heavy Ammonia-ND3

The writers have made an investigation of the infrared absorption spectrum of ND₃, using a rocksalt prism spectrometer. The sample used was very kindly supplied to us by Professor H. S. Taylor, and was prepared from heavy water estimated to be 99 percent pure. The ammonia sample was estimated to be more than 90 percent ND₃, and the absorption spectrum gave no bands characteristic of ordinary ammonia. By following the procedure of Dennison^{1, 2} it is possible to calculate the positions of the fundamental vibrations of this molecule, assuming that the force fields are similar, and that the interatomic distances are the same for the two isotopes. The agreement between observed and calculated values is about as good for the ND₃ as for the values calculated and observed in the literature for the NH₃.³ These are given in Table I. The spectrum should consist of four bands (of the six possible modes of vibration, two are doubly degenerate), in two of which the change in moment should be perpendicular to the symmetry axis. In the NH₃, the bands at 950 cm⁻¹ and 3336 cm⁻¹ have been assigned to the parallel vibrations, and the bands at 1631 cm⁻¹ and 4417 cm⁻¹ are the two fundamental perpendicular vibrations. Of the four bands observed for ND₃, the similarity in the structure of the envelopes of the bands to the corresponding frequencies of NH₃ holds extremely well, and in each case indicates the predicted type of vibration.

TABLE I.

Light ammonia NH ₃				Heavy ammonia ND ₃				
ν	Calc.	Obs.	Type	ν'		Calc.		
ν_3	938	950]			$\nu_3 = 715$	770	
ν_2	1679	1631	L			$\nu_2 = 1189$	1158	Ĩ.
ν_1	3336	3336		$\nu_{1}' =$	0.715	$\nu_1 = 2385$	2421	
ν_4	4466	4417	Ĩ	$\nu_4' =$	0.742	$\nu_4 = 3334$	3287	Ĺ

The writers are continuing this investigation with a grating spectrometer of higher resolution and further details concerning the rotational structure of the bands will be presented in the future.

The writers are much indebted to Professor H. S. Taylor and Dr. H. Eyring and their co-workers at Princeton University for their kindness in supplying the compound.

S. SILVERMAN

J. A. SANDERSON

Physical Laboratory, The Johns Hopkins University, December 2, 1933.

¹ Dennison, Phil. Mag. 1, 195 (1926).

² Dennison, Rev. Mod. Phys. 3, 280 (1931).

³ Schaefer and Matossi, Dasultrarote Spektrum, p. 251.