

ON THE ELEMENTARY ELECTRICAL CHARGE AND THE
AVOGADRO CONSTANT.

BY R. A. MILLIKAN.

I. INTRODUCTORY.

THE experiments herewith reported were undertaken with the view of introducing certain improvements into the oil-drop method¹ of determining e and N and thus obtaining a higher accuracy than had before been possible in the evaluation of these most fundamental constants.

In the original observations by this method such excellent agreement was found between the values of e derived from different measurements (l. c., p. 384) that it was evident that if appreciable errors existed they must be looked for in the constant factors entering into the final formula rather than in inaccuracies in the readings or irregularities in the behavior of the drops. Accordingly a systematic redetermination of all these constants was begun some three years ago. The relative importance of the various factors may be seen from the following review.

As is now well known the oil-drop method rested originally upon the assumption of Stokes's law and gave the charge e on a given drop through the equation

$$e_n = \frac{4}{3} \pi \left(\frac{9\eta}{2} \right)^{\frac{2}{3}} \left(\frac{1}{g(\sigma - \rho)} \right)^{\frac{1}{3}} \frac{(v_1 + v_2)v_1^{\frac{1}{2}}}{F}, \quad (1)$$

in which η is the coefficient of viscosity of air, σ the density of the oil, ρ that of the air, v_1 the speed of descent of the drop under gravity and v_2 its speed of ascent under the influence of an electric field of strength F .

The essential feature of the method consisted in repeatedly changing the charge on a given drop by the capture of ions from the air and in thus obtaining a series of charges with each drop. These charges showed a very exact multiple relationship under all circumstances—a fact which demonstrated very directly the atomic structure of the electric charge. If Stokes's law were correct the greatest common divisor of this series of charges should have been the absolute value of the elementary electrical charge. But the fact that this greatest common divisor failed to come out a constant when drops of different sizes were used showed that Stokes's

¹ R. A. Millikan, *PHYS. REV.*, 32, pp. 349-397, 1911.

law breaks down when the diameter of a drop begins to approach the order of magnitude of the mean free path of a gas molecule. Consequently the following corrected form of Stokes's law for the speed of a drop falling under gravity was suggested.

$$v_1 = \frac{2}{9} \frac{ga^2(\sigma - \rho)}{\eta} \left\{ 1 + A \frac{l}{a} \right\}, \quad (2)$$

in which a is the radius of the drop, l the mean free path of a gas molecule and A an undetermined constant. It is to be particularly emphasized that the term in the brackets was expressly set up merely as a first order correction term in l/a and involved no theoretical assumptions of any sort; further that the constant A was empirically determined through the use only of small values of l/a and that the values of e and N obtained were therefore precisely as trustworthy as were the observations themselves. This fact has been repeatedly overlooked in criticisms of the results of the oil-drop method.¹

Calling then e_1 the greatest common divisor of all the various values of e_n found in a series of observations on a given drop there resulted from the combination of (1) and (2) the equation

$$e \left(1 + A \frac{l}{a} \right)^{\frac{2}{3}} = e_1, \quad (3)$$

or

¹ Indeed M. Jules Roux (Compt. Rendu, 152, p. 1168, May, 1911) has attempted to correct my values of e and N by reducing some observations like mine which he made on droplets of sulphur, with the aid of a purely theoretical value of A which is actually approximately twice too large. The impossibility of the value of A which he assumes he would himself have discovered had he made observations on spheres of different sizes or at different pressures. Such observations whether made on solid spheres or on liquid spheres always yield a value of A about half of that assumed by Roux. Hence his value of e , viz., $e = 4.17 \times 10^{-10}$ rests on no sort of experimental foundation whatsoever. It rests rather on two erroneous assumptions, first the assumption of the correctness of the constants in Cunningham's theoretical equation (Proc. Roy. Soc., 83, p. 357; see also footnote 3, p. 380, PHYS. REV., Vol. 32)—constants which I shall presently show are in no case correct within the limits of experimental error even when inelastic impact is assumed, and second, the assumption that molecules make elastic impact against solid surfaces, an assumption which is completely incorrect as I had already proved by showing that the value of the "slip" term is the same for oil and air as for glass and air (PHYS. REV., Vol. 32, p. 382), which Knudsen also had proved experimentally to be erroneous (Knudsen, An. der Phys., 28, p. 75, 1909, and 35, p. 389) and which for theoretical reasons as well is plainly inadmissible, since were it correct Poiseuille's law could not hold for gases under any circumstances.

But even if Roux had assumed the correct value of A he would still have obtained results several per cent. too low, a fact which must be ascribed either to faulty experimental arrangements or to imperfect knowledge of the density of his sulphur spheres; for solid spheres have been very carefully studied in the Ryerson Laboratory and are in fact found to yield results very close to those obtained with oil drops. Solid spheres however are not nearly so well adapted to a precision measurement of e as are oil drops, since their density and sphericity are always matters of some uncertainty.

$$e = \frac{e_1}{\left(1 + A\frac{l}{a}\right)^{\frac{1}{2}}}. \quad (4)$$

It was from this equation that e was obtained after A had been found by a graphical method which will be more fully considered presently.

The factors then which enter into the determination of e are: (1) The density factor, $\sigma - \rho$; (2) the electric field strength, F ; (3) the viscosity of air, η ; (4) the speeds, v_1 and v_2 ; (5) the drop radius, a ; (6) the correction term constant, A .

Concerning the first two of these factors little need be said unless a question be raised as to whether the density of such minute oil drops might not be a function of the radius. Such a question is conclusively answered in the negative both by theory¹ and by the experiments reported in this paper.

Liquid rather than solid spheres were originally chosen because of the far greater certainty with which their density and sphericity could be known. Nevertheless I originally used liquids of widely different viscosities (light oil, glycerine, mercury) and obtained the same results with them all within the limits of error, thus showing experimentally that so far as this work was concerned, the drops all acted like rigid spheres. More complete proof of this conclusion is furnished both by the following observations and by other careful work on solid spheres soon to be reported in detail by Mr. J. Y. Lee.

The material used for the drops in the following experiments was the highest grade of clock-oil, the density of which, at 23° C., the temperature

¹ The pressure p_2 within an oil drop is given by

$$p_2 = k + \frac{\alpha}{2R}$$

where k is LaPlace's constant of internal pressure, α the constant of surface tension and R the radius. The difference ($p_2 - p_1$) between the pressure within the oil drop and within the oil in bulk is then $\alpha/2R$. But the coefficient of compressibility of a liquid is defined by

$$\beta = \frac{v_2 - v_1}{v_1(p_2 - p_1)}.$$

Now β for oil of this sort never exceeds 70×10^{-6} megadynes per sq. cm. (see Landolt and Bornstein's tables), while α is about 35 dyne cm. R for the smallest drop used (Table XX.) is .00005 cm.; we have then

$$\frac{v_2 - v_1}{v_1} = \beta \frac{\alpha}{2R} = \frac{70 \times 10^{-12} \times 35}{.0001} = .000024.$$

The density of the smallest drop used is then 2 parts in 100,000 greater than that of the oil in bulk. The small drops could then only be appreciably denser than the larger ones if the oil were inhomogeneous and if the atomizing process selected the heavier constituents for the small drops. Such an assumption is negated by the experimental results given in § 9.

at which the experiments were carried out, was found in two determinations made four months apart, to be .9199 with an error of not more than one part in 10,000.

The electric fields were produced by a 5,300-volt storage battery, the P.D. of which dropped on an average 5 or 10 volts during an observation of an hour's duration. The potential readings were taken, just before and just after a set of observations on a given drop, by dividing the bank into 6 parts and reading the P.D. of each part with a 900-volt Kelvin and White electrostatic voltmeter which showed remarkable constancy and could be read easily, in this part of the scale, with an accuracy of about 1 part in 2,000. This instrument was calibrated by comparison with a 750-volt Weston Laboratory Standard Voltmeter certified correct to 1/10 per cent. and actually found to have this accuracy by comparison with an instrument standardized at the Bureau of Standards in Washington. The readings of P.D. should therefore in no case contain an error of more than 1 part in 1,000. As a matter of fact 5,000 volt readings made with the aid of two different calibration curves of the K. & W. instrument made two years apart never differed by more than 1 or 2 parts in 5,000.

The value of F involves in addition to P.D. the distance between the plates, which was as before 16 mm. and correct to about .01 mm. (l. c., p. 351). Nothing more need be said concerning the first two of the above-mentioned factors. The last four however need especial consideration.

2. THE COEFFICIENT OF VISCOSITY OF AIR.

This factor certainly introduces as large an element of uncertainty as inheres anywhere in the oil-drop method. Since it appears in equation (1) in the $3/2$ power an uncertainty of 0.5 per cent. in η means an uncertainty of 0.75 per cent. in e . It was therefore of the utmost importance that η be determined with all possible accuracy. Accordingly two new determinations were begun three years ago in the Ryerson Laboratory, one by Mr. Lachlan Gilchrist and one by Mr. I. M. Rapp. Mr. Gilchrist, whose work has already been published,¹ used a constant deflection method (with concentric cylinders), which it was estimated (l. c., p. 386) ought to reduce the uncertainty in η to 1 or 2 tenths of a per cent. The results have justified this estimate. Mr. Rapp used a form of the capillary tube method which it was thought was better adapted to an *absolute evaluation of η* than have been the capillary tube arrangements which have been commonly used heretofore.² Since Mr.

¹ Lachlan Gilchrist, *PHYS. REV.*, 2d Ser., Vol. 1, p. 124.

² This investigation will shortly be published in full (*PHYS. REV.*, 1913), hence only a bare statement will here be made of the results which are needed for the problem in hand.

Gilchrist completed his work at the University of Toronto, Canada, and Mr. Rapp made his computations and final reductions at Ursinus College, Pa., neither observer had any knowledge of the results obtained by the other. The two results agree within 1 part in 600. Mr. Rapp estimates his maximum uncertainty at 0.1 per cent., Mr. Gilchrist at 0.2 per cent. Mr. Rapp's work was done at 26° C. and gave $\eta_{26} = .00018375$. When this is reduced to 23° C., the temperature used in the following work, by means of formula (5)—a formula¹ which certainly can introduce no appreciable error for the range of temperature here used,—viz.,

$$\eta_t = 0.00018240 - 0.000000493(23 - t); \quad (5)$$

there results

$$\eta_{23} = .00018227.$$

Mr. Gilchrist's work was done at 20.2° C. and gave $\eta_{20.2} = .0001812$. When this is reduced to 23° C. it yields

$$\eta_{23} = .00018257.$$

When this new work, by totally dissimilar methods, is compared with the best existing determinations by still other methods the agreement is exceedingly striking. Thus in 1905, Hogg² made at Harvard very careful observations on the damping of oscillating cylinders and obtained in three experiments at atmospheric pressure $\eta_{23} = 0.0001825$, $\eta_{15.6} = 0.0001790$ and $\eta_{18.6} = 0.0001795$. These last two reduced to 23° C., as above, are 0.0001826 and 0.0001817 respectively and the mean value of the three determinations is

$$\eta_{23} = 0.00018227.$$

Tomlinson's classical determination,³ by far the most reliable of the nineteenth century, yielded when the damping was due primarily to "push" $\eta_{12.65^\circ \text{C.}} = 0.00017746$; when it was due wholly to "drag" $\eta_{11.79^\circ \text{C.}} = 0.00017711$. These values reduced to 15° C., as above, are respectively 0.00017862 and 0.00017867. Hence we may take Tomlinson's direct determination as $\eta_{15} = 0.00017864$. This reduced to 23° C. by Tomlinson's own temperature formula (Holman's) yields $\eta_{23} = 0.00018242$. By the above formula it yields $\eta_{23} = 0.00018256$.

Grindley and Gibson using the tube method on so large a scale⁴ (tube 1/8 inch in diameter and 108 feet long) as to largely eliminate the most

¹ See R. A. Millikan, *Annalen der Physik*, 1913, for a more extended discussion of this and other viscosity formulæ and measurements.

² J. L. Hogg, *Proc. Amer. Acad.*, 40, 18, p. 611, 1905.

³ Tomlinson, *Phil. Trans.*, 177, p. 767, 1886.

⁴ Grindley and Gibson, *Proc. Roy. Soc.*, 80, p. 114, 1908.

fruitful sources of error in this method, namely, the smallness and ununiformity of the bore, obtained at room temperature the following results:¹ $\eta_{25.28^\circ \text{C.}} = .00018347$, $\eta_{23.55^\circ \text{C.}} = .00018241$, $\eta_{12.18^\circ \text{C.}} = .00018257$, and $\eta_{15.4^\circ \text{C.}} = .0001782$. These numbers, reduced to 23°C. as above, are respectively 18,245, 18,241, 18,201, and 18,195. The mean is 18,220. Grindley and Gibson's own formula, $\eta = .0001702 \{1 + .00329t - .0000070t^2\}$, yields $\eta_{23} = .00018245$. We may take then Grindley and Gibson's direct determination as the mean of these two values, viz.: $\eta_{23} = .00018232$.

Collecting then the five most careful determinations of the viscosity of air which so far as I am able to discover have ever been made we obtain the following table.

TABLE I.

Air $\eta_{23} = .00018227$ —Rapp. Capillary tube method. 1913.
Air $\eta_{23} = .00018257$ —Gilchrist. Constant deflection method. 1913.
Air $\eta_{23} = .00018227$ —Hogg. Damping of oscillating cylinder method. 1905.
Air $\eta_{23} = .00018258$ —Tomlinson. Damping of pendular vibrations method. 1886.
Air $\eta_{23} = .00018232$ —Grindley and Gibson. Flow through large pipe method. 1908.
Mean = .00018240

It will be seen, then, that every one of the five different methods which have been used for the absolute determination of η leads to a value which differs by less than 1 part in 1,000 from the above mean value $\eta_{23} = .00018240$. It is surely legitimate then to conclude that the absolute value of η for air is now known with an uncertainty of somewhat less than 1 part in 1,000.²

¹ These numbers represent the reduction to absolute C.G.S. units of all the observations which Grindley and Gibson made between 50°F. and 80°F.

² In obtaining the above mean I have chosen what, after careful study, I have considered to be the best determination by each of the five distinct methods. The transpiration method has been much more commonly used than have the others, and in general, the final result is in good agreement with other careful work by this method. Thus Rankine's final value (Proc. Roy. Soc., A, 83, p. 522, 1910) by a new modification of the capillary tube method, while probably not claiming an accuracy of more than .4 per cent., is, at 10.6°C. , $.0001767$, a value which reduces to $\eta_{23} = .0001828$. Again Fisher's final formula (PHYS. REV., 28, p. 104, 1909) gives $\eta_{23} = .00018218$. Also Holman's much used formula (Phil. Mag., 21, p. 199, 1886, and Tomlinson, Phil. Trans., Vol. 177, part 2, p. 767, 1886) yields $\eta_{23} = .00018237$. In fact the only reliable work on η which I am able to find which is out of line with the value $\eta_{23} = .00018240$ is that by Breiterbach at Leipzig (Ann. der Phys., 5, p. 166, 1901) and that by Schultze (Ann. der Phys., 5, p. 157, 1901) and several other observers at the University of Halle who used Schultze's apparatus (Markowski, Ann. der Phys., 14, p. 742, 1904, and Tanzler, Verh. der D. Phys. Ges., 8, p. 222, 1906). None of these observers however were aiming at an absolute determination, but rather at the effects of temperature and the mixing of gases upon viscosity and their capillaries were too small (of the order .007 cm.) to make possible an absolute determination of high accuracy. Their agreement among themselves upon a value which is about 1.3 per cent. too high is partly accounted for by the fact that several of them used the same tube. None of them made any effort to eliminate the necessarily large error in the measurement of so small a bore (which appears in the result in the fourth power) by taking the mean of η from a considerable number of tubes.

A second question which might be raised in connection with η is as to whether the medium offers precisely the same resistance to the motion through it of a heavily charged drop as to that of an uncharged drop. This question has been carefully studied and definitely answered in the affirmative by the following work (cf. §§ 6 and 10).

3. THE SPEEDS v_1 AND v_2 .

The accuracy previously attained in the measurement of the times of ascent and descent between fixed cross-hairs was altogether satisfactory, but the method which had to be employed for finding the magnifying power of the optical system, *i. e.*, for finding the actual distance of fall of the drop in centimeters, left something to be desired. This optical system was before a short-focus telescope of such depth of focus that it was quite impossible to obtain an accurate measure of the distance between the cross-hairs by simply bringing a standard scale into sharp focus immediately after focusing upon a drop. Accordingly, as stated in the original article, the standard scale was set up at the exact distance from the telescope of the pin-hole through which the drop entered the field. This distance could be measured with great accuracy but *the procedure assumed that the drop remained exactly at this distance throughout the whole of any observation, sometimes of several hours duration.* But if there were the slightest lack of parallelism between gravity and the lines of the electric field the drop would be obliged to drift slowly, and always in the same direction, away from this position, and a drift of 5 mm. was enough to introduce an error of 1 per cent. Such a drift could in no way be noticed by the observer if it took place in the line of sight; for the speeds of the drops were changing *very slowly* anyway because of evaporation, fall in the potential of the battery, etc., and a change in time due to such a drift would be completely masked by other causes of change. This source of uncertainty was well recognized at the time of the earlier observations and steps were taken at the beginning of the present work to eliminate it. It was in fact responsible for an error of nearly two per cent.

A new optical system was built, consisting of an achromatic objective of 28 mm. aperture and 12.5 cm. focal length and an eyepiece of 12 mm. focal length. The whole system was mounted in a support which could be moved bodily back and forth by means of a horizontal screw of $\frac{1}{2}$ mm. pitch. In an observation the objective was 25 cm. distant from the drop, which was kept continually in sharp focus by advancing or withdrawing the whole telescope system. The depth of focus was so small that a motion of $\frac{1}{2}$ mm. blurred badly the image of the drop. The eyepiece

was provided with a scale having 80 horizontal divisions and the distance between the extreme divisions of this scale (the distance of fall in the following experiments) could be regularly duplicated with an accuracy of at least 1 part in 1,000, by bringing a standard scale (Société Genevoise) into sharp focus. (The optical path when the scale was viewed was made exactly the same as when the drop was viewed.) The distance of fall, then, one of the most uncertain factors of the preceding determination, was now known with at least this degree of precision.

The accuracy of the *time* determinations can be judged from the data in Tables IV.-XIX. On account of the great convenience of a direct-reading instrument these time measurements were all made, not with a chronograph, as heretofore, but with a Hipp chronoscope which read to 0.002 second. This instrument was calibrated by comparison with the standard Ryerson Laboratory clock under precisely the same conditions as those under which it was used in the observations themselves and found to have an error between 0 and 0.2 per cent. depending upon the time interval measured. For the sake of enabling others to check all the computations herein contained if desired, as well as for the sake of showing what sort of consistency was attained in the measurement of time intervals there are given in Table II. the calibration readings for the 30 sec. interval and in Table III. the results of similar readings for all the intervals used.

TABLE II.

TABLE III.

Chronoscope Readings for 30 Sec. Interval.		Clock Interval, Sec.	Chronoscope Interval.	Corr'n Applied, Per Cent.
29.962	29.990	6	6.0146	-0.26
29.988	29.958	10	10.0018	0.00
29.986	29.920	16	16.0080	0.00
29.930	29.972	20	19.9835	+0.07
29.964	29.976	30	29.9695	+0.10
30.002	30.006	40	39.9436	+0.14
29.940	29.979	60	59.9072	+0.16
29.998	30.018	114	113.795	+0.20
29.930	29.926	120	119.782	+0.20
29.967	29.972			
Corr'n = +.1 per cent.				

The change in the per cent. correction with the time interval employed is due to the difference in the reaction times of the magnet and spring contact at make (beginning) and at break (end). All errors of this sort are obviously completely eliminated by making the calibration observations under precisely the same conditions as the observations on the drop. In Tables IV. to XIX. the recorded times are the uncorrected chronograph

readings. The corrections are obtained by interpolation in the last column of Table III.

Under the head of possible uncertainties in the velocity determinations are to be mentioned also the effects of a distortion of the drop by the electric field. Such a distortion would increase the surface of the drop, and hence the speed imparted to it per dyne of *electric* force would not be the same as the speed imparted per dyne of gravitational force when the field was off and the drop had the spherical form. The following observations were made in such a way as to bring to light such an effect if it were of sufficient magnitude to exert any influence whatever upon the accuracy of the determination of e by this method (cf. §§ 6 and 10).

Similarly objection has been made to the oil-drop method on the ground that, on account of internal convection, fluid drops would not move through air with the same speed as solid drops of like diameter and mass. Such objection is theoretically unjustifiable in the case of oil drops of the sizes here considered.¹ Nevertheless the experimental demonstration of its invalidity is perhaps worth while and is therefore furnished below.

4. THE RADIUS "a."

The radius of the drop enters only into the correction term (see equation 4) and so long as this is small need not be determined with a high degree of precision. It is most easily obtained by the following procedure which differs slightly from that originally employed (l. c., p. 379).

It will be seen that the equation (l. c., p. 353)

$$\frac{v_1}{v_2} = \frac{mg}{Fe - mg} \quad (6)$$

contains no assumption whatever save that a given body moves through a given medium with a speed which is proportional to the force acting upon it. Substitution in this equation of $m = \frac{4}{3}\pi a^3(\sigma - \rho)$ and the solution of the resulting equation for a gives

$$a = \sqrt[3]{\frac{3Fe}{4\pi g(\sigma - \rho)} \frac{v_1}{(v_1 + v_2)}} \quad (7)$$

The substitution in this equation of an approximately correct value of e yields a with an error but one third as great as that contained in the assumed value of e . The radius of the drop can then be determined from (7) with a very high degree of precision as e becomes more and more accurately known. In the following work the value of e substituted in (7) to obtain a was 4.78×10^{-10} but the final value of e obtained would

¹ Hadamard, Compt. Rendus, 1911.

not have been appreciably different if the value substituted in (7) to obtain a had been 5 per cent. or 6 per cent. in error. The determination of a therefore introduces no perceptible error into the evaluation of e .

5. THE CORRECTION-TERM CONSTANT A .

This constant was before graphically determined (l. c., p. 379) by plotting the values of e_1^3 as ordinates and those of l/a as abscissæ and observing that if we let $x = l/a$, $y = e_1^3$ and $y_0 = e^3$ equation (3) may be written in the form

$$y_0(1 + Ax) = y \quad (8)$$

or

$$A = \frac{\frac{dy}{dx}}{y_0} = \frac{\text{slope}}{y \text{ intercept}}. \quad (9)$$

Now even if the slope were correctly determined by the former observations all of the above-mentioned sources of error would enter into the value of the intercept and hence would modify the value of A .

As a matter of fact however the accuracy with which the slope itself was determined could be much improved, for with the preceding arrangement it was necessary to make all the observations at atmospheric pressure and the only way of varying l/a was by varying a , *i. e.*, by using drops of different radii. But when a was very small the drops moved exceedingly slowly under gravity and the minutest of residual convection currents produced relatively large errors in the observed speeds, *i. e.*, in e_1 . If for example the time of fall over a distance of 2 mm. is 20 minutes it obviously requires an extraordinary degree of stagnancy to prevent a drift in that time of say .2 mm. due to convection. But this would introduce an error of 10 per cent. into e_1 . Furthermore with these slow drops Brownian movements introduce errors which can only be eliminated by taking a very large number of readings¹ and this is not in general feasible with such drops. It is quite impossible then by working at a single pressure to obtain from the graph mentioned above a line long enough (l. c., p. 379) to make the determination of its slope a matter of great precision. Accordingly in the new observations the variation of l/a was effected chiefly through the variation of l , *i. e.*, of the pressure p , rather than of a . This made possible not only the accurate evaluation of e , but also the solution of the interesting question as to the law of fall of a given drop through air at reduced pressures.

¹ Fletcher, *PHYS. REV.*, 33, p. 92, 1911.

6. METHOD OF TESTING THE ASSUMPTIONS INVOLVED IN THE OIL-DROP METHOD.

In order to make clear the method of treatment of the following observations a brief consideration of the assumptions underlying the oil-drop method must here be made. These assumptions may be stated thus:

1. The drag which the medium exerts upon a given drop is unaffected by its charge.
2. Under the conditions of observation the oil drops move through the medium essentially as would solid spheres. This assumption may be split into two parts and stated thus: Neither (2a) distortions due to the electric field nor (2b) internal convection within the drop modify appreciably the law of motion of an oil drop.¹
3. The density of oil droplets is independent of their radius down to $a = .00005$ cm.

Of these assumptions (2a) is the one which needs the most careful experimental test.² It will be seen that it is contained in the fundamental equation of the method (see (7)) which may be written in the form

$$e_n = \frac{mg}{Fv_1}(v_1 + v_2). \quad (10)$$

Or still more conveniently in the form

$$e_n = \frac{mgt_g}{F} \left(\frac{1}{t_g} + \frac{1}{t_F} \right), \quad (11)$$

in which t_g and t_F are the respective time intervals required by the drop to fall under gravity and to rise under the field F the distance between the cross-hairs.

In order to see how the assumption under consideration can be tested let us write the corresponding equation after the same drop has caught n' additional units, namely,

$$e_{n+n'} = \frac{mgt_g}{F} \left(\frac{1}{t_g} + \frac{1}{t_{F'}} \right). \quad (12)$$

The subtraction of (11) from (12) gives

$$e_{n'} = \frac{mgt_g}{F} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right). \quad (13)$$

¹M. Brillouin has in addition suggested (see p. 149, La Théorie du Rayonnement et les Quanta) that the drops may be distorted by the molecular bombardment, but Einstein's reply (l. c., p. 150) to this suggestion is altogether unanswerable, and, in addition, such a distortion, if it existed, would make the value of e given by the oil-drop method *too small* instead of too large.

²Professor Lunn has however subjected it to a theoretical study and has in this way demonstrated its validity (PHYS. REV., XXXV., p. 227, 1912).

Now equations (11) and (12) show, since $mg t_0/F$ remains constant, that as the drop changes charge the successive values of its charge are proportional to the successive values assumed by the quantity $(1/t_0 + 1/t_F)$ and the elementary charge itself is obviously this same constant factor $mg t_0/F$ multiplied by *the greatest common divisor* of all these successive values. It is to be observed too that since $1/t_0$ is in these experiments generally large compared to $1/t_F$ the value of this greatest common divisor, which will be denoted by $(1/t_0 + 1/t_F)_0$, is determined primarily by the time of fall *under gravity*, and is but little affected by the time in the field. On the other hand equation (13) shows that the greatest common divisor of the various values of $(1/t_{F'} - 1/t_F)$, which will be designated by $(1/t_{F'} - 1/t_F)_0$, when multiplied by the same constant factor $mg t_0/F$, is also the elementary electrical charge. In other words $(1/t_0 + 1/t_F)_0$ and $(1/t_{F'} - 1/t_F)_0$ are one and the same quantity, but while the first represents essentially a speed measurement when the field is off, the second represents a speed measurement in a powerful electric field. If then the assumption under consideration is correct we have two independent ways of obtaining the quantity which when multiplied by the constant factor $mg t_0/F$ is the elementary electrical charge, but if on the other hand the distortion of the drop by the field modifies the law of motion of the oil drop through the medium then $(1/t_0 + 1/t_F)_0$ and $(1/t_{F'} - 1/t_F)_0$ will not be the same. Now *a very careful experimental study of the relations of $(1/t_0 + 1/t_F)_0$ and $(1/t_{F'} - 1/t_F)_0$ shows so perfect agreement that no effect of distortion in changing measurably the value of e can be admitted.*¹ (See Tables IV. to XIX.)

Turning next to assumption (1), this can be tested in three ways, all of which have been tried with negative results. First a drop containing from one up to six or seven elementary charges can be completely discharged and its time of fall under gravity when uncharged compared with its time when charged. Second, the multiple relationships shown in the successive charges carried by a given drop may be very carefully examined. They cannot hold exactly if when the drop is heavily charged it suffers a larger drag from the medium than when it is lightly charged. Third, when drops having widely different charges and different masses are

¹ It may be pointed out in passing that the above discussion brings to light a method of obtaining e which is independent of a viscosity measurement; for $(1/t_{F'} - 1/t_F)_0$ can be obtained for a body which is heavy enough to be weighed upon a micro-balance. Such a body would fall so rapidly that $1/t_0$ could not be measured, but it could be computed from the measurement of $1/t_{F'}$ and $1/t_F$ and the equation $(1/t_0 + 1/t_{F'})_0 = (1/t_{F'} - 1/t_F)_0$. Either (12) or (13) could then be solved for e after m had been determined by direct weighing. A consideration of the sources of error in this method shows however that it cannot be made as accurate as the present method which involves the coefficient of viscosity of air.

brought to the same value of l/a by varying the pressure, the value of e_1 (which is proportional to $(v_1 + v_2)_0$), should come out smaller for the heavily than for the lightly charged drops. The following observations show that this is not the case.

The last criterion is also a test for (2b) for if internal convection modifies the speed of fall of a drop as Perrin wishes to assume that it may,¹ it must play a smaller and smaller rôle as the drop diminishes in size, hence varying l/a by diminishing a cannot be equivalent to varying l/a by increasing l . In other words the value of e_1 obtained from work on a large drop at a low pressure should be different from that obtained from work on a small drop at so high a pressure that l/a has the same value as for the large drop.

Finally if the density of a small drop is greater than that of a large one (see assumption 3) then, for a given value of l/a , the small drop will show a larger value of e_1 than the large one inasmuch as the computation of e_1 is based on a constant value of σ . *The fact, then, that for a given value of l/a the value of e_1 actually comes out independent of the radius or charge of the drops shows conclusively either that no one of these possible sources of error exists, or else that they neutralize one another so that for the purposes of this experiment they do not exist.* That they do not exist at all is shown by the independent theoretical and experimental tests mentioned above. This removes I think every criticism which has been suggested of the oil-drop method of determining e and N .

7. SUMMARY OF IMPROVEMENTS IN METHOD.

In order to obtain the consistency shown in the following observations it was found necessary to take much more elaborate precautions to suppress convection currents in the air of the observing chamber than had at first been thought needful.

To recapitulate, then, the improvements which have been introduced into the oil-drop method, consist in (1) a redetermination of η ; (2) an improved optical system; (3) an arrangement for observing speeds at all pressures; (4) the more perfect elimination of convection; (5) the experimental proof of the correctness of all the assumptions underlying the method, viz., (a) that a charge does not alter the drag of the medium on the charged body; (b) that the oil drops act essentially like solid spheres; (c) that the density of the oil drops is the same as the density of the oil in bulk.

¹ La Théorie du Rayonnement et les Quanta, p. 239—Rapports et Discussions de la Réunion tenue à Bruxelles, Novembre, 1911. Edited by Langevin and de Broglie. Gauthier-Villars.

8. THE EXPERIMENTAL ARRANGEMENTS.

The experimental arrangements are shown in Fig. 1. The brass vessel *D* was built for work at all pressures up to 15 atmospheres but since the present observations have to do only with pressures from 76 cm. down these were measured with a very carefully made mercury manometer *m* which at atmospheric pressure gave precisely the same reading as a

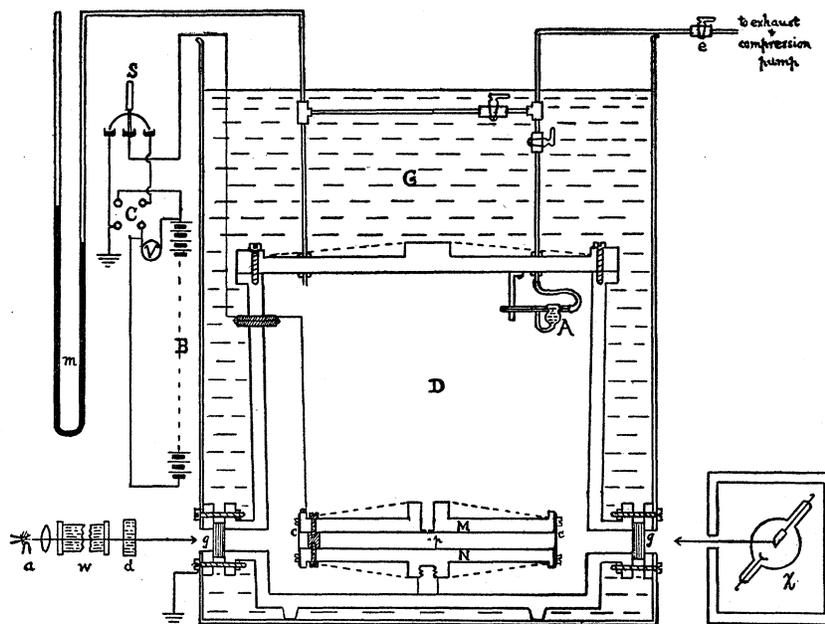


Fig. 1.

standard barometer. Complete stagnancy of the air between the condenser plates *M* and *N* was attained first by absorbing all of the heat rays from the arc *A* by means of a water cell *w*, 80 cm. long, and a cupric chloride cell¹ *d*, and second by immersing the whole vessel *D* in a constant temperature bath *G* of gas-engine oil (40 liters) which permitted, in general, fluctuations of not more than $.02^{\circ}$ C. during an observation. This constant temperature bath was found essential if such consistency of measurement as is shown below was to be obtained. A long search for causes of slight irregularity revealed nothing so important as this and after the bath was installed all of the irregularities vanished. The atomizer *A* was blown by means of a puff of carefully dried and dust-free air introduced through the cock *e*. The air about the drop *p* was ionized

¹ See Coblenz, Bulletin of the Bureau of Standards, Washington, D. C., Vol. 7, p. 660, 1911.

when desired by means of Röntgen rays from X which readily passed through the glass window g . To the three windows g (two only are shown) in the brass vessel D correspond, of course, three windows in the ebonite strip c which encircles the condenser plates M and N . Through the third of these windows, set at an angle of about 18° from the line Xpa and in the same horizontal plane, the oil drop is observed.

9. THE OBSERVATIONS.

The record of a typical set of readings on a given drop is shown in Table IV. The first column, headed t_g , gives the successive readings on the time of descent under gravity. The fourth column, headed t_F , gives the successive times of ascent under the influence of the field F as measured on the Hipp chronoscope. These two columns contain all the data which is used in the computations. But in order to have a test of the stagnancy of the air a number of readings were also made with a stop-watch on the times of ascent through the first half and through the whole distance of ascent. These readings are found in the second and third columns, the times for the first half under the head $\frac{1}{2}t_F$, the times for the whole distance under the head t_F . It will be seen from these readings that there is no indication whatever of convection, since the readings for the one half distance have uniformly one half of the value of the readings for the whole distance, within the limits of error of a stop-watch measurement. This sort of a test was made on the majority of the drops, but since no further use is made of these stop-watch readings they will not be given in succeeding tables.

The fifth column, headed $1/t_F$, contains the reciprocals of the values in the fourth column after the correction found from Tables II. and III. has been applied. The sixth column contains the successive differences in the values of $1/t_F$ resulting from the capture of ions. The seventh column, headed n' , contains the number of elementary units caught at each change, a number determined simply by observing by what number the quantity just before it in column 6 must be divided to obtain the constancy shown in the eighth column, which contains the successive determinations of $(1/t_F' - 1/t_F)_0$ (see § 6). Similarly the ninth column, headed n , gives the total number of units of charge on the drop, a number determined precisely as in the case of the numbers in the seventh column by observing by what numbers the successive values of $(1/t_g + 1/t_F)$ must be divided to obtain the constancy shown in the tenth column, which contains the successive values of $(1/t_g + 1/t_F)_0$. Since n' is always a small number and in some of the changes almost always has the value 1 or 2 its determination for any change is obviously never a matter of the slightest

uncertainty. On the other hand, n is often a large number, but with the aid of the known values of n' it can always be found with absolute certainty so long as it does not exceed say 100 or 150. It will be seen from the means at the bottom of the eighth and the tenth columns that in the case of this drop the two ways discussed in § 6 of obtaining the number which when multiplied by $mg\tau_0/F$ is the elementary electrical charge yield absolutely identical results.

TABLE IV.

Drop No. 6.

t_g Sec.	$\frac{1}{2}F_s$ Sec.	t_{F_s} Sec.	t_{F_c} Sec.	$\frac{1}{t_F}$	$\left(\frac{1}{t_{F'}} - \frac{1}{t_F}\right)$	n'	$\frac{1}{n'}\left(\frac{1}{t_{F'}} - \frac{1}{t_F}\right)$	n	$\frac{1}{n}\left(\frac{1}{t_g} + \frac{1}{t_{F'}}\right)$
11.848	39.9	80.2	80.708	.01236	.03234	6	.005390	18	.005366
11.890	11.2	22.4	22.366	.04470					
11.908			22.390	.04470	.03751	7	.005358	24	.005371
11.904	11.2	22.4	22.368	.007192					
11.882	70.6	140.4	140.565	.007192	.005348	1	.005348	17	.005375
11.906	39.9	79.6	79.600	.01254					
11.838			34.748	.02870	.01616	3	.005387	21	.005376
11.816			34.762	.02870					
11.776			34.846	.03414	.026872	5	.005375	22	.005379
11.840			29.286	.03414					
11.904	14.6	29.3	29.236	.007268	.021572	4	.005393	17	.005380
11.870	69.3	137.4	137.308	.007268					
11.952	17.6	34.9	34.638	.02884	.01623	3	.005410	21	.005382
11.860				.01623					
11.846			22.104	.04507	.04307	8	.005384	24	.005386
11.912			22.268	.04507					
11.910			500.1	.002000	.04879	9	.005421	16	.005387
11.918			19.704	.05079					
11.870			19.668	.05079	.03874	7	.005401	25	.005399
11.888			77.630	.01285					
11.894	38.9	77.6	77.806	.01285	.01079	2	.005395	18	.005390
11.878	21.0	42.6	42.302	.02364					
11.880					Means		.005386		.005384

Duration of exp. = 45 min.,
 Plate distance = 16 mm.,
 Fall distance = 10.21 mm.,
 Initial volts = 5,088.8.
 Final volts = 5,081.2.
 Temperature = 22.82° C.,
 Pressure = 75.62 cm.,
 Oil density = .9199,
 Air viscosity¹ = $1,824 \times 10^{-7}$,
 Radius (a) = .000276 cm.,
 l/a = .034,
 Speed of fall = .08584 cm./sec.,
 $e_1 = 4.991 \times 10^{-10}$.

¹ In the above and in all the following tables the computations were made on the basis of the assumption $\eta_{23} = 1,825 \times 10^{-7}$ instead of $\eta_{23} = 1,824 \times 10^{-7}$ (see § 2). The reduction to the latter value has been made only in the final value of e (see § 10).

TABLE V.

Drop No. 16.

t_g	t_F	$\frac{I}{t_F}$	n'	$\frac{I}{n'} \left(\frac{1}{t_F} - \frac{1}{t_g} \right)$	n	$\frac{I}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.638							
18.686							
18.689	17.756	.05628	5	.006908	16	.006853	$V_i = 5106$ $V_f = 5100$ $t = 23.7^\circ \text{ C.}$ $p = 74.68$ $v_1 = .05449$ $a = .0002188$ $l/a = .04390$ $e_1 = 5.065$
18.730	17.778				.02174	3	
18.686	45.978	5	.006860	13			
18.726	45.870			.021826	4	.006825	
18.772	45.716	.001441	2				
18.740	45.758			.03574	12	.006856	
18.724	694.0	.008439	2				
18.720	27.95			.02217	9	.006876	
18.816	118.388	.02890	11				
18.816	45.030			.02227	2	.006889	
18.716	34.564	.008518	2				
18.804	44.826			.022295	2	.006889	
18.746	117.198	.022295	2				
18.746	44.784			.022295	2	.006889	
18.790							
18.738				.006860		.006861	

TABLE VI.

Drop No. 14.

t_g	t_F	$\frac{I}{t_F}$	n'	$\frac{I}{n'} \left(\frac{1}{t_F} - \frac{1}{t_g} \right)$	n	$\frac{I}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.606							
18.732							
18.784							
18.700	46.172	.02163	5	.006874	11	.006820	$V_i = 5077$ $V_f = 5073$ $t = 23.09^\circ \text{ C.}$ $p = 75.28$ $v_1 = .05451$ $a = .0002185$ $l/a = .04348$ $e_1 = 5.064$
18.730	17.896	.05600			5	.006886	
18.652	17.818		.02157	1			
18.656	46.328	.01484			6	.006840	
18.730	46.258		.05588	17			
18.760	46.266	.06305			9	.006853	
18.708	67.473		.001370	6			
18.658	67.148	.04266			4	.006850	
18.668	67.148		.01526	1			
18.826	17.896	.008389			1	.006784	
18.710	15.868		.001605	1			
18.802	15.854	.001370			6	.006882	
18.778	730.0		.04266	4			
18.790	23.376	.01526			1	.006871	
18.846	23.504		.008389	1			
18.804	65.416	.001605			1	.006784	
18.662	118.970		.001605	1			.006784
18.704	622.8	.001605			1	.006784	
18.730							.006850

In order to show the sort of consistency attained in this work the complete records are given in Tables V., VI., VII., VIII. and IX. of five drops of practically the same size, taken at very nearly the same pressures. These are the only drops of this size which were observed with the present arrangement so that they represent the regular run of observations rather than a selected group. The symbols in the last column V_i , V_f , t , p , v_1 , and a stand for initial volts, final volts, temperature, pressure, velocity under gravity, and radius, respectively. It will be seen from the second columns (see Table V.) that in general in spite of the precautions taken against evaporation (l. c., p. 388) the drops do evaporate very slowly, since with a given charge the speed in the field very slowly increases. It is for this reason that the numbers in the next to the last column tend to increase very slightly. This slow change introduces no error into e_1 , provided corresponding values of t_g , t_F and V (volts) are combined to obtain e_1 . The mean values taken throughout the same time-interval

TABLE VII.
Drop No. 13.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.626						
18.716						
18.602	21.206	5	.006696	15	.006716	$V_i = 5030$ $V_f = 5018$ $t = 23.0^\circ \text{C.}$ $p = 76.06$ $v_1 = .05463$ $a = .0002191$ $l/a = .04290$ $e_1 = 5.043$
18.634	21.110					
18.694	72.590					
18.678	36.454					
18.598	48.870					
18.612	142.838	2	.006735	9	.006722	
18.622	72.428	1	.006792	10	.006729	
18.795	36.722	2	.006732			
18.732	36.652			12	.006729	
18.780	36.602					
18.736	140.891	3	.006719	9	.006732	
18.754	48.302	2	.006799			
18.712	48.266	4	.006756	11	.006745	
18.760	20.942			15	.006747	
18.716						
18.740						
18.695			.006753		.006731	

constitute such corresponding values. On account of this slow change, however, the readings from which the differences $(1/t_{F'} - 1/t_F)$ are taken should be separated by as small a time interval as possible. It is for this reason that in Table V., for example, to obtain the difference $(1/t_{F'} - 1/t_F)$ corresponding to the change from $t_F = 45$ sec. to $t_F = 694$ sec.

TABLE VIII.

Drop No. 15.

t_g	t_F	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.908				
18.948	14.348	.006790	.006801	$V_i = 5085$ $V_f = 5071$ $t = 23.83^\circ \text{ C.}$ $p = 75.24$ $v_1 = .05274$ $a = .0002182$ $l/a = .04368$ $e_1 = 5.040$
18.958	14.352			
18.950	45.070			
18.904	34.338			
18.988	34.432	.006807	.006816	
18.896	34.334			
19.094	44.864	.006858	.00620	
18.936	34.430			
18.980	548.0	.006831	.006831	
19.030	34.14			
18.872	44.654	.006766	.006834	
18.808	44.520			
18.910	44.526			
18.946	64.28			
18.936	114.404	.006751	.006835	
18.950	113.452			
19.030	63.934	.006728	.006832	
19.116	23.260			
19.018	23.350			
18.959		.006817	.006823	

TABLE IX.

Drop No. 17.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.402						
18.556						
18.350	14.966	6	.006727	18	.006752	$V_i = 5065$ $V_f = 5055$ $t = 23.06^\circ \text{ C.}$ $p = 73.47$ $v_1 = .05545$ $a = .0002209$ $e/a = .04411$ $e_1 = 5.054 \times 10^{-10}$
18.470	14.870					
18.412	37.376					
18.318	37.594					
18.388	37.368	1	.006780	12	.006748	
18.376	50.228					
18.406	76.154	1	.006780	11	.006749	
18.392	155.6					
18.492		1	.006682	10	.006742	
18.415						
			.006682	9	.006749	
18.415			.006742		.006748	

the last two of the 45 sec. readings are averaged instead of the whole four. It is worthy of note, too, in this connection that if a change in the time in the field takes place first from 45 sec. to 694 sec. and then im-

TABLE X.

Drop No. 42.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
18.360						
18.362						
18.328	45.378	6	.007027	11	.006960	$V_i = 4240$ $V_f = 4236$ $t = 22.94^\circ \text{C}$ $p = 15.72 \text{ cm.}$ $v_1 = .05564 \frac{\text{cm.}}{\text{sec.}}$ $a = .0002060 \text{ cm.}$ $l/a = .2210$ $e_1 = 6.244 \times 10^{-10}$
18.444	45.208			17	.006983	
18.442	15.574			16	.006983	
18.330	17.468			10	.006965	
18.330	66.000	6	.007011	9	.006962.	
18.428	66.020					
18.174	122.600	1	.006993	11	.006960	
18.388	45.414	2	.006951			
18.392	45.282	2	.006911	9	.006971	
18.240	45.110					
18.314	45.282					
18.324	121.600					
18.312	121.200					
18.390						
18.347			.006978		.006969	

TABLE XI.

Drop No. 46.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
25.856	10.334	10	.007086	19	.007121	$V_i = 3406$ $V_f = 3396$ $t = 22.81^\circ \text{C.}$ $p = 14.68 \text{ cm.}$ $v_1 = .03937 \text{ cm.}$ $a = .0001690 \frac{\text{cm.}}{\text{sec.}}$ $l/a = .2886$ $e_1 = 6.719 \times 10^{-10}$
25.946	10.342			9	.007159	
25.890	38.604			1	.007195	
25.876	53.450			3	.007152	
25.818	24.906	3	.007183	11	.007152	
25.934	30.248			10	.007172	
25.956	30.016			13	.007174	
25.876	18.282			14	.007159	
26.064	16.216	2	.007155	11	.007161	
26.018	24.860			12	.007168	
25.902	21.062	2	.007155	10	.007171	
25.860	30.126					
25.930	20.884					
25.994	21.130	5	.007170	7	.007167	
25.814	21.174					
26.012	85.922					
25.904	226.600					
25.802	86.368	1	.007148	6	.007159	
25.918	86.406					
25.816						
25.909			.007160		.007161	

mediately back again to the 45 second value the mean of the two values of $(1/t_{F'} - 1/t_F)$ thus obtained is independent of any error arising from the evaporation of the drop. For this reason the mean value of the quantity $1/n'(1/t_{F'} - 1/t_F)$ is in general much more trustworthy than might be inferred from the variation in the individual numbers from which this mean is obtained.

TABLE XII.

Drop No. 53.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
33.432	28.494	2	.006501	10	.006491	$V_i = 2463$ $V_f = 2440.2$ $t = 23.16^\circ \text{ C.}$ $p = 12.61$ $v_i = .03055$ $a = .0001439$ $l/a = .3945$ $e_1 = 7.450 \times 10^{-10}$
33.346	28.624			12	.006493	
33.172	20.806			9	.006491	
33.310	20.832	3	.006503	10	.006491	
33.380	35.032			6	.006482	
33.306	28.548	4	.006505	5	.006485	
33.346	111.244			6	.006475	
33.328	400.000	1	.006471	6	.006475	
33.684	111.706	1	.006436			
33.484						
33.379			.006484		.006487	

TABLE XIII.

Drop No. 48.

t_g	t_F	$\frac{1}{t_F}$	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
32.366	20.330	.04919	5	.006164	13	.006157	$V_i = 2561.$ $V_f = 2557.$ $t = 22.81^\circ \text{ C.}$ $p = 15.35 \text{ cm.}$ $v_i = .03150$ $a = .0001501 \text{ cm.}$ $l/a = .3104$ $e_1 = 6.866 \times 10^{-10}$
32.398	23.140	.04313			12	.006165	
32.217	23.230				7	.006166	
32.342	80.786	.01231			7	.006137	
32.358	81.416				14	.006144	
32.438	18.113	.05517	6	.006142	8	.006147	
32.412	18.144		7	.006156			
32.236	54.490	.01832	1	.006080			
32.424	81.486	.01224					
32.356				.006131		.006156	

Nevertheless in practically all of the following work in view of the large number of observations in the t_g column, the mean at the bottom of the column $1/n(1/t_g + 1/t_F)$ is considered more trustworthy than the mean at the bottom of the column $1/n'(1/t_{F'} - 1/t_F)$ and it has been in fact exclusively used in the computation of e_1 . Only in the case of 2 or 3 of the fastest drops (Tables XVI. and XVII.) are the uncertainties in

TABLE XIV.

Drop No. 47.

t_g	t_F	$\frac{I}{t_F}$	n'	$\frac{I}{n'} \left(\frac{I}{t_F} - \frac{I}{t_F} \right)$	n	$\frac{I}{n} \left(\frac{I}{t_g} + \frac{I}{t_F} \right)$
12.836	11.960	.08369	13	.003812	42	.003840
12.994	11.938					
12.892	10.704					
12.914	29.350	.03411	5	.003858	29	.003851
12.930	29.280					
12.902	29.108					
12.838	32.866	.03034	1	.003860	28	.003854
12.808	33.068					
12.920	32.894					
12.886	32.892	.01108	2	.003898	23	.003855
12.908	29.204					
12.902	32.950					
12.916	90.040	.01496	2	.003878	24	.003854
12.850	90.152					
12.854	66.762					
12.912	66.650	.00716	2	.003852	22	.003852
12.908	139.8					
12.870	139.4					
12.882	66.942	.01492	24	.003855		
12.891				.003861		.003852

$V_i = 2555.$
 $V_f = 2545.$
 $t = 22.83^\circ \text{C.}$
 $p = 9.70 \text{ cm.}$
 $v_1 = .07921 \text{ cm./sec.}$
 $a = .0002380 \text{ cm.}$
 $l/a = .3097$
 $e_1 = 6.841 \times 10^{-10}$

TABLE XV.

Drop No. 41.

t_g	t_F	$\frac{I}{t_F}$	n'	$\frac{I}{n'} \left(\frac{I}{t_F} - \frac{I}{t_F} \right)$	n	$\frac{I}{n} \left(\frac{I}{t_g} + \frac{I}{t_F} \right)$
24.016	42.188	.02369	1	.009380	8	.009336
24.142	42.078					
24.130	42.098					
24.070	69.900	.01431	1	.009389	6	.009328
24.000	203.200	.004921	4	.009255	5	.009316
24.030	23.844	.04194	9	.009286	9	.009286
24.046	30.606	.02326	1	.009260	8	.009289
24.028	42.800					
23.968	42.944					
24.018	71.400	.01400	7	.009276	7	.009276
23.770		.03259	2	.009295	6	.009277
23.882	30.652					
24.008				.009314		.009301

$V_i = 5065$
 $V_f = 5059$
 $t = 23.05^\circ \text{C.}$
 $p = 19.01 \text{ cm.}$
 $v_1 = .04253$
 $a = .0001816$
 $l/a = .1394$
 $e_1 = 6.097 \times 10^{-10}$

$(I/t_g + I/t_F)_0$ greater perhaps than those in $(I/t_F' - I/t_F)_0$ and in these cases the two were so near together that there was no object in using the latter instead of the former. It should be stated that all time intervals

longer than 150 sec. were measured with the stop-watch rather than with the chronoscope and are not subjected to the corrections in Tables II. and III. In general, too, only differences in $t_F' - t_F$ amounting to as much as 20 seconds are used for the determination of $(1/t_F' - 1/t_F)$ since obviously the observational error is large when $t_F' - t_F$ is small.

TABLE XVI.

Drop No. 1.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_F'} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$	
	17.514	11	.003244	89	.003235	$V_i = 5174$ $V_f = 5162$ $t = 23.00^\circ \text{ C.}$ $p = 75.80 \text{ cm.}$ $v_1 = .23565 \frac{\text{cm.}}{\text{sec.}}$ $a = .0005856$ $l/a = .01615$ $e_1 = 4.877 \times 10^{-10}$
	17.500			78	.003233	
	46.724			10	.003216	
4.346	46.494	77	.003228			
4.400	18.722	11	.003256	86	.003227	
4.310	18.616					
4.376	56.280	9	.003217	77	.003219	
4.370	56.276					
4.422	21.362					
4.392	21.476	10	.003199	87	.003216	
4.290	58.936					
4.392	58.194					
4.360	58.456					
4.346	58.904					
4.368	20.412					
4.350	20.404					
4.363	20.322	101	.003211			
corrected to	10.710					
	10.660					
	20.290					
	20.296					
4.328 ¹			.003226		.0032245	

A study of Tables IV. to IX. shows, *first*, a striking agreement between the values of $(1/t_F' - 1/t_F)_0$ and $(1/t_g + 1/t_F)_0$ the largest divergence being found in Table V. where it amounts to 0.3 per cent., and *second*, a satisfactory agreement between the values of e_1 obtained from the different

¹ These corrected times in Tables XVI. and XVII. are obtained by a slight extrapolation upon a calibration curve plotted from Table III. These two drops are the largest used and fall too rapidly for the most accurate work, but were introduced so as to test as thoroughly as possible the assumptions that even the largest of these drops fall like rigid spheres and that heavily charged drops experience no more drag from the medium than do uncharged ones. The drop of Table XVII. is evaporating more rapidly than any other used (see column t_F), probably because of the low pressure (4.46 cm.), hence the fourth column numbers are here obtained only from immediately adjacent observations in the second column. Each of these fourth column numbers shows the influence of the evaporation, but it is interesting to see how it is eliminated from the mean.

drops of approximately the same size, the largest divergence from the mean value being here too 0.3 per cent.

TABLE XVII.

Drop No. 56.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$			
5.092	19.480							
5.010	19.356			136	.001837			
5.082	19.374							
5.044	19.202							
—	29.214							
	29.022			127	.001832			
5.044	28.686							
5.030	38.688							
	38.620	7	.001786	122	.001839	$V_i = 1701$ $V_f = 1699$ $t = 23.21^\circ \text{ C.}$ $p = 4.46 \text{ cm.}$ $v_1 = .20256$ $a = .0003653 \text{ cm.}$ $l/a = .4396$ $e_1 = 7.777 \times 10^{-10}$		
5.066	75.136							
5.120	73.752			117	.001811			
5.046	72.758	9	.001837					
5.128	33.034						126	.001814
4.912	32.788						123	.001814
5.036	39.956			121	.001814			
5.080	46.790	5	.001796					
5.122	80.820						116	.001816
5.058	80.152	3	.001877					
cor'd to	55.206					119	.001810	
	34.398	6	.001823	125	.001818			
5.039 ¹			.001824		.001821			

TABLE XVIII.

Drop No. 22.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{1}{t_F} \right)$			
40.518	17.760	4	.01012	8	.01012	$V_i = 5009$ $V_f = 5001$ $t = 23.22^\circ \text{ C.}$ $p = 76.42 \text{ cm.}$ $v_1 = .02515$ $a = .0001475$ $l/a = .06350$ $e_1 = 5.168 \times 10^{-10}$		
40.522	63.100						3	.01015
40.382								
40.542								
40.650	15.138			9	.01009			
40.696	27.716			6	.01011			
40.478	15.060	4	.01011	9				
40.458	38.606						5	.01012
40.494	38.382	2	.01008					
40.802	172.626							.01013
40.542			.01010		.01012			

A few typical records of observations made at different pressures (see "p," last column) and on drops of different sizes (see "a," last column) are presented in Tables X to XIX.

TABLE XIX.

Drop No. 52.

t_g	t_F	n'	$\frac{1}{n'} \left(\frac{1}{t_{F'}} - \frac{1}{t_F} \right)$	n	$\frac{1}{n} \left(\frac{1}{t_g} + \frac{2}{t_F} \right)$	
50.364	44.446 } 30.254 } 30.326 } 22.904 }	1	.010515	4	.010566	$V_i = 3362$
50.442				5	.010556	$V_f = 3340$
50.416				6	.010575	$t = 22.98$
50.770				5	.010532	$p = 16.95$
50.178				3	.010573	$v_1 = .020216$
50.456				2	.01047	$a = .0001183$
50.460	83.716 } 796.000 }	1	.010664	3	.010573	$l/a = .3568$
50.118				2	.010528	$e_1 = 7.210 \times 10^{-10}$
50.400			.01055		.010555	

Table XX. contains a complete summary of the results obtained on all of the 58 different drops upon which complete series of observations like the above were made during a period of 60 consecutive days. It will be seen from this table that these observations represent a 30-fold variation in l/a (from .016, drop No. 1, to .444, drop No. 58), a 17-fold variation in p (from 4.46 cm., drop No. 56, to 76.27 cm., drop No. 10), a 12-fold variation in a (from 4.69×10^{-5} cm., drop No. 28, to 58.56×10^{-5} cm., drop No. 1) and a variation in the number of free electrons carried by the drop from 1 on drop No. 28 to 136 on drop No. 56. The time of fall of drop No. 28 was also tested when it was completely discharged, as have been the times of many other drops which carried most of the time but one electron.

Much larger variations both in a and p , and therefore in l/a , might have been used, and have in fact been used, for finding the law of fall of a drop through rarefied air, but for the end here sought, namely, the most accurate possible determination of e , it was found desirable to keep the t_g interval for the most part between the limits 10 sec. and 40 sec., in order to avoid chronograph errors on the one hand and Brownian movement irregularities on the other. That neither of these sources of error is appreciable in these observations may be seen from a study of Tables IV.-XIX., which are thoroughly representative of the work on all the drops.

10. RESULTS AND DISCUSSION.

It will be seen at once from equation (4) that the value of e is simply the value of e_1 for which $l/a = 0$, so that if successive values of $e_1^{\frac{2}{3}}$ are plotted as abscissæ and of l/a as ordinates the intercept of the resulting curve on the $e_1^{\frac{2}{3}}$ axis is $e^{\frac{2}{3}}$. Furthermore if A is a constant then the curve in question is a straight line and A is the slope of this line divided by the

TABLE XX.

No.	Tem. °C.	P.D. (Volts).	t_g (Secs.).	V_1 cm./sec.	$(V_1 + V_2)/b$	n	$\alpha \times 10^6$ cm.	β (cm. Hg.).	$r/\rho a$	l/a	$e_1 \times 10^{10}$	$e_1^2 \times 10^8$	$e_2^2 \times 10^8$
1	23.00	5.168	4.363	.2357	.003293	77-102	58.56	75.80	22.52	.01615	4.877	61.90	61.14
2	22.80	5.120	8.492	.1202	.004670	27-36	32.64	75.00	40.85	.02933	4.981	62.82	61.26
3	23.46	5.100	9.905	.1032	.004996	22-27	30.29	73.71	44.88	.03212	4.971	62.75	61.04
4	22.85	5.163	10.758	.09489	.005211	18-36	28.94	75.20	45.92	.03288	5.001	63.00	61.24
5	23.08	5.072	10.663	.09575	.005176	20-30	29.14	73.25	46.85	.03353	4.982	62.82	61.13
6	22.82	5.085	11.880	.08584	.005497	17-24	27.54	75.62	48.11	.03437	4.991	62.93	61.09
7	23.79	5.090	11.950	.08368	.005480	19-22	27.57	75.10	48.44	.03466	4.981	62.82	61.07
8	23.50	5.158	12.540	.08141	.005623	16-19	26.90	75.30	49.52	.03544	5.016	63.12	61.23
9	22.87	5.139	13.562	.07375	.005962	19-23	25.71	75.00	51.73	.03702	5.016	63.13	61.15
10	23.25	5.015	15.380	.06641	.006174	13-22	24.31	76.27	54.09	.03871	5.010	63.08	61.02
11	23.01	5.066	15.193	.06720	.006087	11-14	24.36	73.90	55.52	.03973	5.015	63.12	61.00
12	23.00	5.080	15.985	.06375	.006416	12-16	23.70	75.14	56.15	.04018	5.028	63.24	61.10
13	23.00	5.024	15.695	.05463	.006873	9-15	21.91	76.06	59.94	.04290	5.043	63.35	61.06
14	23.09	5.077	18.730	.05451	.006988	8-16	21.85	75.28	60.78	.04348	5.064	63.53	61.21
15	23.85	5.078	18.959	.05274	.006966	8-18	21.78	75.24	61.03	.04368	5.040	63.33	61.07
16	23.70	5.103	18.738	.05449	.007005	9-16	21.87	74.68	61.33	.04390	5.065	63.54	61.21
17	23.06	5.060	18.415	.05545	.006890	9-18	22.06	73.47	61.69	.04411	5.054	63.43	61.00
18	22.83	5.093	26.130	.03907	.008339	5-13	18.45	75.54	71.74	.05134	5.098	63.82	61.08
19	22.95	5.033	28.568	.03570	.008651	5-9	17.63	75.87	74.77	.05350	5.120	64.00	61.12
20	23.00	5.094	9.480	.10772	.005058	23-32	30.54	41.77	78.40	.05612	5.145	64.22	61.23
21	23.08	5.018	35.253	.02893	.009660	4-11	15.80	74.32	85.08	.06089	5.166	64.36	61.11
22	23.22	5.005	40.542	.02515	.010332	3-9	14.75	76.42	88.70	.06350	5.168	64.40	61.01
23	22.76	5.098	39.900	.02554	.010510	3-6	14.85	75.40	89.35	.06395	5.190	64.59	61.18
24	23.16	5.050	12.466	.08189	.005896	15-28	26.44	37.19	101.8	.07283	5.269	65.24	61.35
25	22.98	5.066	15.157	.06737	.006399	12-17	24.01	38.95	107.2	.07660	5.278	65.28	61.20
26	23.20	4.572	7.875	.12980	.004324	33-40	33.07	24.33	124.4	.08892	5.379	66.06	61.31
27	23.18	4.570	9.408	.1085	.004730	23-29	30.23	25.37	130.4	.09330	5.381	66.16	61.18
28	23.00	5.145	84.270	.1211	.01595	1-4	4.69	75.83	130.3	.09322	5.379	66.14	61.16
29	22.99	5.073	23.223	.04393	.008488	6-12	19.06	33.47	156.8	.1117	5.529	67.36	61.37

Mean = 61.120

Mean = 61.138

No.	Temp. ° C.	P.D. (Volts).	t_p (Sec.).	V_1 cm./sec.	$(V_1 + V_2)_0$	n	$a \times 10^6$ cm.	ρ (cm. HG).	$1/\rho a$	l/a	$e_1 \times 10^{10}$	$e_1^2 \times 10^8$	$e_1^3 \times 10^6$	$e_1^3 \times 10^8$
30	23.19	5.090	26.830	.03801	.009111	5-10	17.77	35.18	160.2	.1147	5.507	67.18	61.06	
31	22.89	5.098	38.479	.02649	.011180	3-5	14.71	36.51	176.5	.1263	5.621	68.12	61.38	
32	23.06	5.070	14.060	.07246	.006762	12-17	24.29	21.12	195.0	.1394	5.692	68.67	61.22	
33	23.07	4.582	18.229	.05601	.006981	10-13	21.33	23.86	196.6	.1405	5.687	68.64	61.13	
34	23.06	5.061	38.010	.02682	.011205	3-8	14.72	34.01	199.8	.1429	5.714	68.84	61.20	
35	23.00	4.246	9.265	.11032	.004653	27-34	29.84	16.00	209.5	.1499	5.739	69.07	61.07	
36	22.91	4.236	9.879	.10340	.004863	24-28	28.74	15.67	222.0	.1589	5.820	69.71	61.23	
37	23.06	4.236	12.040	.08496	.005362	18-24	26.27	16.75	227.5	.1625	5.821	69.72	61.03	
38	22.94	2.556	10.657	.09581	.003109	32-43	27.49	14.70	247.5	.1771	5.935	70.61	61.16	
39	23.00	5.054	19.950	.05115	.008370	8-15	20.12	19.73	251.8	.1802	5.910	70.41	60.79	
40	23.09	5.058	21.130	.04830	.008865	7-9	18.38	18.54	278.3	.1993	6.076	71.72	61.09	
41	23.05	5.062	24.008	.04254	.009496	6-8	18.16	19.01	289.6	.2073	6.110	72.03	60.97	
42	22.94	4.238	18.347	.05564	.007110	9-17	20.60	15.72	308.8	.2210	6.224	73.04	61.24	
43	23.18	3.254	13.909	.07340	.004729	16-28	23.70	13.55	311.0	.2227	6.214	72.83	60.95	
44	23.04	4.231	29.114	.03503	.009273	5-9	16.16	17.17	360.6	.2379	6.466	74.77	61.00	
45	22.97	3.317	29.776	.03425	.007430	5-12	15.90	17.27	364.2	.2606	6.537	75.30	61.39	
46	22.81	3.401	25.909	.03937	.007311	6-19	16.90	14.68	403.3	.2886	6.719	76.71	61.30	
47	22.83	2.550	12.891	.07921	.003935	18-42	23.80	9.70	432.8	.3097	6.841	77.66	61.13	
48	22.80	2.559	32.326	.03150	.006286	7-14	15.01	15.35	433.8	.3104	6.866	77.85	61.28	
49	23.02	3.370	14.983	.06815	.011353	8-9	22.00	10.10	448.8	.3221	6.936	78.36	61.22	
50	23.45	2.535	11.659	.08757	.003783	25-30	24.88	8.60	466.7	.3340	6.978	78.67	60.85	
51	23.48	2.539	10.924	.09346	.003615	27-34	25.69	8.26	470.7	.3368	7.024	79.02	61.04	
52	22.98	3.351	50.400	.02021	.010775	2-6	11.83	16.95	498.5	.3568	7.210	80.40	61.36	
53	23.16	2.451	33.379	.03055	.006623	5-10	14.39	12.61	551.3	.3945	7.470	82.19	61.13	
54	23.46	2.533	19.227	.05347	.005314	11-17	18.87	9.03	587.8	.4112	7.661	83.73	61.18	
55	22.90	2.546	24.254	.04206	.006041	9-18	16.72	10.11	591.5	.4233	7.672	83.82	61.22	
56	23.21	1.700	5.058	.20256	.001861	117-136	36.53	4.46	614.2	.4396	7.777	84.57	61.11	
57	23.12	2.321	15.473	.06599	.004360	18-24	20.85	7.74	619.7	.4435	7.774	84.54	60.87	
58	23.03	3.388.5	24.33	.04196	.008183	6-10	16.62	9.070	620.2	.4439	7.810	84.83	61.14	

Mean = 61.138

Mean of all numbers in last column = 61.138
Mean of first 23 numbers = 61.120

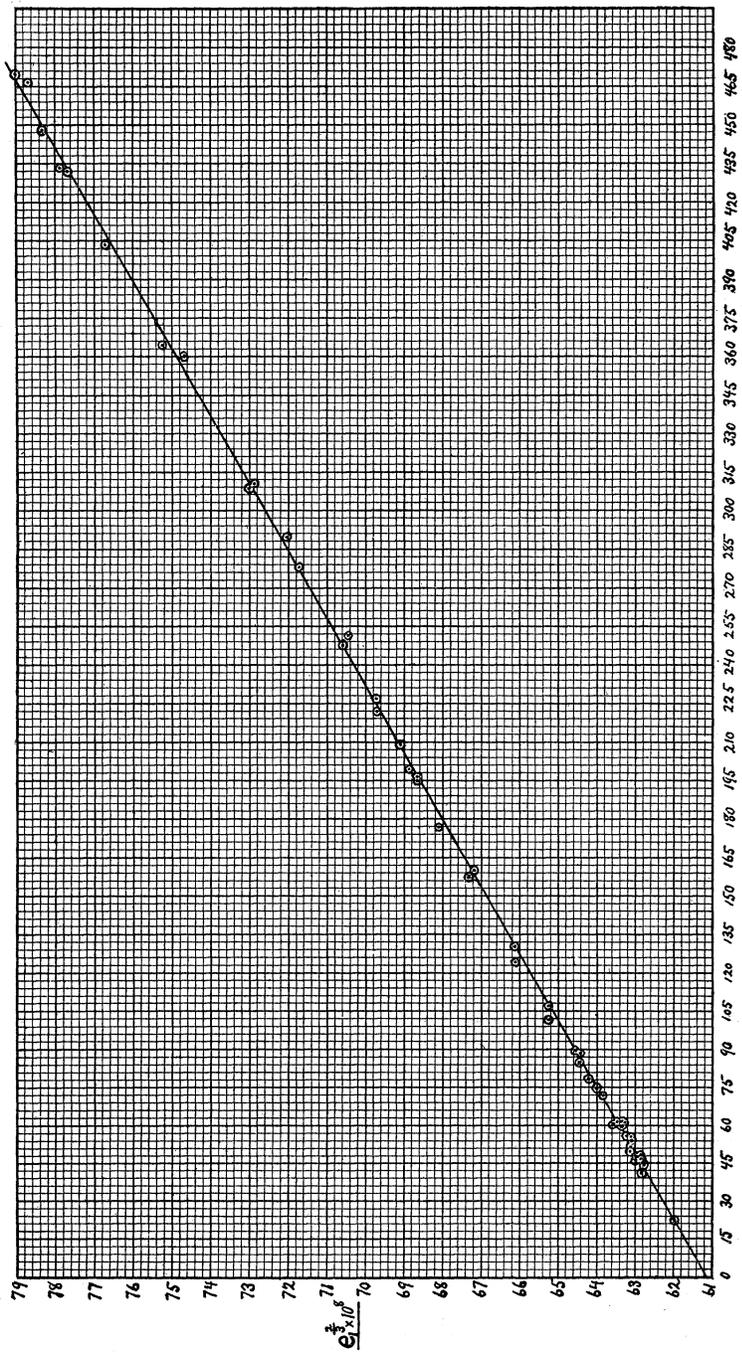
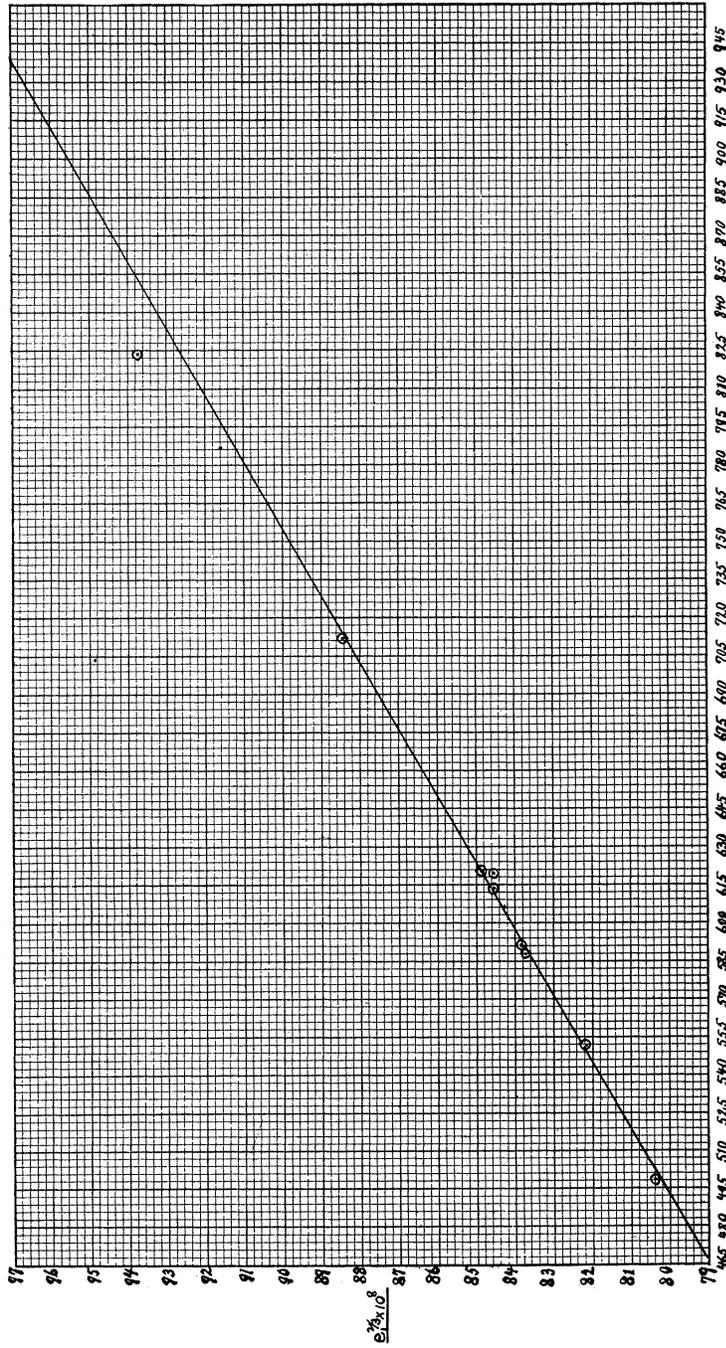


Fig. 2.



$\frac{1}{ka}$
Fig. 3.

y intercept (see equation 9). In view of the uncertainty in l due to the fact that k in the equation $\eta = knm\bar{c}l$ has never been exactly evaluated, it was thought preferable to write the correction term to Stokes's law (see (2 and 3) in the form $(1 + b/pa)^{-1}$ instead of in the form $(1 + Al/a)^{-1}$ and then to plot $e_1^{\frac{2}{3}}$ against $1/pa$. Nevertheless in view of the greater ease of visualization of l/a all the values of this quantity corresponding to successive values of $1/pa$ are given in Table XX., k being taken, merely for the purposes of this computation, as .3502 (Boltzmann). Fig. 2 shows the graph obtained by plotting the values of $e_1^{\frac{2}{3}}$ against $1/pa$ for the first 51 drops of Table XX., and Fig. 3 shows the extension of this graph to twice as large values of $1/pa$ and $e_1^{\frac{2}{3}}$. It will be seen that there is not the slightest indication of a departure from a linear relation between $e_1^{\frac{2}{3}}$ and $1/pa$ up to the value $1/pa = 620.2$ which corresponds to a value of l/a of .4439 (see drop No. 58, Table XX.). Furthermore the scale used in the plotting is such that a point which is one division above or below the line in Fig. 2 represents in the mean an error of 2 in 700. *It will be seen from Figs. 2 and 3 that there is but one drop in the 58 whose departure from the line amounts to as much as 0.5 per cent. It is to be remarked, too, that this is not a selected group of drops but represents all of the drops experimented upon during 60 consecutive days, during which time the apparatus was taken down several times and set up anew. It is certain then that an equation of the form (2) holds very accurately up to $l/a = .4$. The last drop in Fig. 3 seems to indicate the beginning of a departure from this linear relationship. Since such departure has no bearing upon the evaluation of e , discussion of it will be postponed to another paper.*

Attention may also be called to the completeness of the answers furnished by Figs. 2 and 3 to the questions raised in § 6. Thus drops No. 27 and 28 have practically identical values of $1/pa$ but while No. 28 carries, during part of the time, but 1 unit of charge (see Table XX.) drop No. 27 carries 29 times as much and it has about 7 times as large a diameter. Now if the small drop were denser than the large one (see assumption 3, § 6) or if the drag of the median upon the heavily charged drop were greater than its drag upon the one lightly charged (see assumption 1, § 6), then for both these reasons drop 27 would move more slowly relatively to drop 28 than would otherwise be the case and hence $e_1^{\frac{2}{3}}$ for 27 would fall below $e_1^{\frac{2}{3}}$ for drop 28. Instead of this the two $e_1^{\frac{2}{3}}$'s fall so nearly together that it is impossible to represent them on the present scale by two separate dots. Drops 52 and 56 furnish an even more striking confirmation of the same conclusion, for both drops have about the same value for l/a and both are exactly on the line though No. 56

carries at one time 68 times as heavy a charge as No. 52 and has three times as large a radius. In general the fact that Figs. 2 and 3 show no tendency whatever on the part of either the very small or the very large drops to fall above or below the line is experimental proof of the joint correctness of assumptions 1, 3, and $2b$ of § 6. The correctness of $2a$ was shown by the agreement throughout Tables IV. to XIX. between $1/n'(1/t_F' - 1/t_F)$ and $1/n(1/t_0 + 1/t_F)$.

The values of e^3 and b obtained graphically from the y -intercept and the slope in Fig. 2 are $e^3 = 61.13 \times 10^{-8}$ and $b = .0006254$, p being measured, for the purposes of Fig. 2 and of this computation in mm. of Hg at 23° C. and a being measured in cm. The value of A (equations 2 and 3) corresponding to this value of B is .874 instead of .817 as originally found. Cunningham's theory gives, in terms of the constants here used, $A = 788$.¹

Instead however of taking the result of this graphical evaluation of e^3 it is more accurate to reduce each of the observations on e_1^3 to e^3 by means of the above value of B and the equation

$$e^3 \left(1 + \frac{b}{pa} \right) = e_1^3. \quad (14)$$

The results of this reduction are contained in the last column of Table XX. These results illustrate very clearly the sort of consistency obtained in these observations. *The largest departure from the mean value found anywhere in the table amounts to 0.5 per cent., and "the probable error" of the final mean value computed in the usual way is 16 in 61,000.*

Instead however of using this final mean value as the most reliable evaluation of e^3 it was thought preferable to make a considerable number of observations at atmospheric pressure on drops small enough to make t_0 determinable with great accuracy and yet large enough so that the whole correction term to Stokes's law amounted to but a few per cent., since in this case, even, though there might be a considerable error in the correction-term constant b , such error would influence the final value of e by an inappreciable amount. The first 23 drops of Table XX. represent such observations. It will be seen that they show slightly greater consistency than do the remaining drops in the table and that the correction-term reductions for these drops all lie between 1.3 per cent. (drop No. 1) and 5.6 per cent. (drop No. 23) so that even though b were in error by as much as 3 per cent. (its error is actually not more than .5 per cent.) e^3 would be influenced by that fact to the extent of but 0.1 per cent. The mean value of e^3 obtained from the first 23 drops is 61.12×10^{-8} , a

¹ PHYS. REV., 32, p. 380; also footnote.

number which differs by 1 part in 3,400 from the mean obtained from all the drops.

When correction is made for the fact that the numbers in Table XX. were obtained on the basis of the assumption $\eta_{23} = .0001825$, instead of $\eta_{23} = .0001824$ (see § 2) the final mean value of e^3 obtained from the first 23 drops is 61.085×10^{-8} . This corresponds to

$$e = 4.774 \times 10^{-10} \text{ electrostatic units.}$$

Since the value of the Faraday constant has now been fixed virtually by international agreement¹ at 9,650 absolute electromagnetic units and since this is the number N of molecules in a gram molecule times the elementary electrical charge, we have

$$N \times 4.774 \times 10^{-10} = 9,650 \times 2.9990 \times 10^{10};$$

$$\therefore N = 6.062 \times 10^{23}.$$

Although the probable error in this number computed by the method of least squares from Table XX. is but one part in 3,000 it would be erroneous to infer that e and N are now known with that degree of precision, for there are four constant factors entering into all of the results in Table XX. and introducing uncertainties as follows. The coefficient of viscosity η which appears in the $3/2$ power introduces into e and N a maximum possible uncertainty of 0.1 per cent. The distance between the condenser plates (16.00 mm.) is correct to .01 mm., and therefore, since it appears in the 1st power in e , introduces a maximum possible error of something less than 0.1 per cent. The voltmeter readings have a maximum possible error of rather less than 0.1 per cent., and carry this in the 1st power into e and N . The cross-hair distance which is uniformly duplicatable to one part in a thousand appears in the $3/2$ power and introduces an uncertainty of no more than 0.1 per cent. The other factors introduce errors which are negligible in comparison. *The uncertainty in e and N is then that due to 4 continuous factors each of which introduces a maximum possible uncertainty of 0.1 per cent.* Following the usual procedure we may estimate the uncertainty in e and N as the square root of the sum of the squares of these four uncertainties, that is, as 2 parts in 1,000. We have then finally:

$$e = 4.774 \pm .009 \times 10^{-10}$$

and

$$N = 6.062 \pm .012 \times 10^{23}.$$

The difference between these numbers and those originally found by the oil-drop method, viz., $e = 4.891$ and $N = 5.992$ is due to the fact

¹ Atomic weight of silver 107.88. Electrochemical equivalent of silver 0.01118.

that this much more elaborate and prolonged study has had the effect of changing every one of the three factors η , A , and d (= cross-hair distance) in such a way as to lower e and to raise N . The chief change however has been due to the elimination of the faults of the original optical system.

II. COMPARISON WITH OTHER MEASUREMENTS.

So far as I am aware, there is at present no determination of e or N by any other method which does not involve an uncertainty at least 15 times as great as that represented in the above measurements.

Thus the *radioactive method* yields in the hands of Regener¹ a count of the α particles which gives e with an uncertainty which he estimates at 3 per cent. This is as high a precision I think as has yet been claimed for any α particle count,² though Geiger and Rutherford's photographic registration³ method will doubtless be able to improve it.

The *Brownian Movement* method yields results which fluctuate between Perrin's value⁴ $e = 4.24 \times 10^{-10}$, and Fletcher's value,⁵ 5.01×10^{-10} , with Svedberg's measurements⁶ yielding the intermediate number 4.7×10^{-10} .

The *radiation method* of Planck⁷ yields N as a product of $(c_2)^3$ and σ . The latest Reichsanstalt value of c_2 is 1.436⁸ while Coblenz,⁹ as the result of extraordinarily careful and prolonged measurements obtains 1.4456. The difference in these two values of $(c_2)^3$ is 2 per cent. Westphal¹⁰ estimates his error in the measurement of σ at .5 per cent. though reliable observers differ in it by 5 per cent. or 6 per cent. We may take then 3 per cent. as the limit of accuracy thus far attained in measurements of e or N by other methods. *The mean results by each one of the three other methods fall well within this limit of the value found above by the oil-drop method.*

12. COMPUTATION OF OTHER FUNDAMENTAL CONSTANTS.

For the sake of comparison and reference, the following fundamental constants are recomputed on the basis of the above measurements:

¹ Regener, Sitz. Ber. d. k. Preuss. Acad., 37, p. 948, 1909.

² Rutherford and Geiger, Proc. Roy. Soc., 81, p. 155, 1908.

³ Geiger and Rutherford, Phil. Mag., 24, p. 618, 1912.

⁴ J. Perrin, C. R., 152, p. 1165, 1911.

⁵ H. Fletcher, Phys. Rev., 33, p. 107, 1911.

⁶ Svedberg, Arkiv f. Kemi, etc., utg. af K. Sv. Vetensk. Akad., 2, 29, 1906. See also Svedberg, "Die Existenz der Molekule," p. 136. Leipzig, 1912.

⁷ Planck, Vorlesungen über die Theorie der Wärmestrahlung, 2d edition, 1913, p. 166.

⁸ See Planck, Vorles., p. 163.

⁹ Coblenz, Journal of the Washington Academy of Sciences, Vol. 3, p. 178, April, 1913.

¹⁰ Wm. H. Westphal, Verh. d. D. Phys. Ges., 13, p. 987, Dec., 1912.

1. The number n of molecules in 1 c.c. of an ideal gas at 0° 76 is given by

$$n = \frac{N}{V} = \frac{6.062 \times 10^{23}}{22,412} = 2.705 \times 10^{19}.$$

2. The mean kinetic energy of agitation E_0 of a molecule at 0° C. is given by

$$pV = \frac{1}{3}Nmu^2 = \frac{2}{3}NE_0 = RT,$$

$$\therefore E_0 = \frac{3}{2} \frac{p_0 V_0}{N} = \frac{3 \times 1,013,700 \times 22,412}{2 \times 6.062 \times 10^{23}} = 5.621 \times 10^{-14} \text{ ergs.}$$

3. The constant ϵ of molecular energy defined by $E_0 = \epsilon T$ is given by

$$\epsilon = \frac{E_0}{T} = \frac{5.621 \times 10^{-14}}{273.11} = 2.058 \times 10^{-16} \frac{\text{ergs}}{\text{degrees}}.$$

4. The Boltzmann entropy constant k defined by $S = k \log W$ is given by¹

$$k = \frac{R}{N} = \frac{p_0 V_0}{TN} = \frac{2}{3}\epsilon = 1.372 \times 10^{-16} \frac{\text{ergs}}{\text{degrees}}.$$

All of these constants are known with precisely the accuracy attained in the measurement of e .

5. The Planck "Wirkungsquantum" h can probably be obtained considerably more accurately as follows than in any other way. From equation 292, p. 166, of the "Wärmestrahlung," we obtain²

$$h = \frac{k^{\frac{3}{2}}}{c} \left(\frac{48\pi\alpha}{a} \right)^{\frac{1}{2}} = \frac{(1.372 \times 10^{-16})^{\frac{3}{2}}}{2.999 \times 10^{10}} \cdot \left(\frac{48\pi \cdot 1.0823}{7.39 \times 10^{-15}} \right)^{\frac{1}{2}} = 6.620 \times 10^{-27}$$

which gives h with the same accuracy attainable in the measurement of $k^{\frac{3}{2}}/a$ in which a is the Stefan-Boltzmann constant. If Westphal's estimate of his error in the measurement of this constant is correct, viz., 0.5 per cent., it would introduce an uncertainty of but 0.2 per cent. into h . This is about that introduced by the above determination of $k^{\frac{3}{2}}$, hence the above value of h should not be in error by more than 0.4 per cent.

6. The constant c_2 of the Wien-Planck radiation law may also be computed with much precision from the above measurements. For also from equation 292 of the "Wärmestrahlung" we obtain

$$c_2 = \left(\frac{48\pi\alpha k}{a} \right)^{\frac{1}{2}} = \left(\frac{48\pi \cdot 1.0823 \cdot 1.372 \times 10^{-16}}{7.39 \times 10^{-15}} \right)^{\frac{1}{2}} = 1.4470 \text{ cm. degrees.}$$

¹ See Planck's Vorles., p. 129.

² c = velocity of light, α = a numerical factor, and $a = 4\sigma/c$. Westphal's value of σ is 5.54×10^{-8} which corresponds to $a = 7.39 \times 10^{-15}$.

Since both k and a here appear in the $1/3$ power, the error in c should be no more than 0.2 per cent., provided Westphal's error is no more than 0.5 per cent.

The difference between this and Coblenz's mean value, viz., 1.4456 is but 0.1 per cent. The agreement is then entirely satisfactory. A further independent check is found in the fact that Day and Sosman's location of the melting point of platinum at 1755° C.¹ is equivalent to a value of $c_2 = 1.4475$.² On the other hand, the last Reichsanstalt value of c_2 , viz., 1.437, is too low to fit well with these and Westphal's measurements. It fits perfectly however with a combination of the above value of e and Shakespear's³ value of σ , viz., $\sigma = 5.67 \times 10^{-5}$.

13. SUMMARY.

The results of this work may be summarized in the following table in which the numbers in the error column represent in the case of the first six numbers estimated limits of uncertainty rather than the so-called "probable errors" which would be much smaller. The last two constants however involve Westphal's measurements and estimates and Planck's equations as well as my own observations.

TABLE XXI.

Elementary electrical charge.....	$e = 4.774 \pm .009 \times 10^{-10}$
Number of molecules per gram molecule.....	$N = 6.062 \pm .012 \times 10^{23}$
Number of gas molecules per c.c. at 0° 76.....	$n = 2.705 \pm .005 \times 10^{19}$
Kinetic energy of a molecule at 0° C.....	$E_0 = 5.621 \pm .010 \times 10^{-14}$
Constant of molecular energy.....	$\epsilon = 2.058 \pm .004 \times 10^{-16}$
Constant of the entropy equation.....	$k = 1.372 \pm .002 \times 10^{-16}$
Elementary "Wirkungsquantum".....	$h = 6.620 \pm .025 \times 10^{-27}$
Constant of the Wien displacement law.....	$c_2 = 1.4470 \pm 0030$

I take pleasure in acknowledging the able assistance of Mr. J. Yinbong Lee in making some of the above observations. Mr. Lee has also repeated with my apparatus the observations on oil at atmospheric pressure with results which are nearly as consistent as the above. Using my value of b he obtains, as a mean of measurements on 14 drops, a value of e which differs from the above by less than 1 part in 6,000, although its probable error computed as in the case of Table XX. is 1 part in 2,000.

RYERSON PHYSICAL LABORATORY,
UNIVERSITY OF CHICAGO,
June 2, 1913.

¹ Amer. Jour. Sci., 30, p. 3, 1910.

² Coblenz, Journal of the Washington Academy of Sciences, Vol. 3, p. 13.

³ G. A. Shakespear, Proc. Roy. Soc., 86, 180, 1911.